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(g) $\rightleftharpoons 2\text{HI}(g)$ Some nitrogen gas is added No change; N_2 is not a component of this reaction system. (s) + $\text{HCl}(g)$ reaction is carried out in an open container Because HCl is a gas that can escape from the system, the reaction is forced to the right. This is the basis for the commercial production of hydrochloric acid.

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