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Section 10 1
Measuring Matter

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CHAPTER

We use the term the Mole to represent Avogadro's number. A mole of atoms is 6.022×10^{23} atoms, much like a dozen atoms is 12 atoms. AMU of Aluminum = 26.98 AMU of Copper is =63.55 AMU of Carbon is 12.01. So, 26.98 grams of aluminum, 63.55 grams of copper, and 12.01 grams of carbon each contain 6.022×10^{23} atoms!

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Composition 225 The
most recent
measurement of
Avogadro's number
shows that it is
 $6.02214199 \times 10^{23}$

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units/mole. In this book, the measurement is rounded to 6.022×10^{23} units/mol. Avogadro's number is used to count any kind of particle, as shown in Figure 1. The Mole Is a Counting Unit

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The mole is a convenient unit for expressing the relative amounts of substances in chemical reactions. The burning of carbon in oxygen can be written with the 2 oxygen atoms bonded in a single O_2 molecule: The mole is the most common unit used to express the quantity of a chemical substance.

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Stoichiometry Study
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Define the following: a.

Stoichiometry-the
study of the
quantitative

relationships between
the amounts of
reactants used and the
products formed by a
chemical reaction. b.

Mole -The SI unit used
to measure the amount

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of a substance that
contains 6.02×10^{23}
atoms of that
substance.

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make products which
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with negatively
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mole of a pure
substance. Avogadro's
number. ... Atomic
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The electron

configuration of
hydrogen is like that of
Group 1 metals. 4.

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1. pair 2. Study Guide.

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Mole: A measurement
of matter. What are 1.
How many moles is
1.50 x 10²³ molecules
of NH₃? 2.49 x ...

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MOLES/
STOICHIOMETRY Mole-
a unit of measurement

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that expresses the amount of atoms, molecules or some other unit. The number of items in one mole is commonly referred to Avogadro's number which equals 6.022×10^{23} . Example: One mole of carbon equals 6.022×10^{23} atoms or particles of carbon.

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